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6.27 Descriptive

Abbreviations (S) Solid

(l) Liquid (g) Gas (aq

) aqueous Determining

State at Room

Temperature □ All

metals are solids except

for mercury, which is a

liquid.

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bond. ionic bond.

cation. anion. forms

when 2 or more atoms

rearrange valence

electrons to incr \square . forms

when valence electrons

are transferred from one

atom to a \square . atom loses

electrons to become

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positively charged. atom
gains electrons to
become negatively
charged.

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Acharya Quantum

Numbers N, l, m, s

Used to describe an

electron in an atom n

Principle quantum

number Represents

main energy level of

electron maximum # of

electrons in an Energy

level = $2n^2$ Example:

What is the maximum

number of electrons that

can be in the 5th main

energy level? 1 The 2 nd

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quantum number l

Describes the orbital
shape within an energy
level l number of orbital
shapes possible in
energy level = n Orbital
Shapes designated s, p ...

note_ep304aditi.pdf -
orientation of orbital in
space s ...

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give away electrons,

Answers.
ions are positively

charged, ions.

nonmetals. take

electrons, ions are

negatively charged,

anions. writing ionic

formulas. 1. write the

symbols for the element.

-the cation is written

first. -The anion is

written second. 2.

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Determine the charge on
each ion.

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Answers

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Nayan Jacquet

International System

(SI) Measuring system

used in science AKA

metric system Length -

distance between two

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points Base unit- meter

(M) Instrument- meter
stick or metric ruler

Mass -amount of matter
in an object. Base unit-

gram (G) Instrument-
balance Volume

-amount of space an
object occupies Base
unit-Liter (L)

Instrument- metric ruler
or graduated cylinder

Time - interval between
two occurrences Base

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Solutions: A Special
Type of Mixture ...
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Name _____

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7.1 Lab results: 1 doz
grains of rice = _____g
(Use this fact as a
conversion

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